





**THE CHEMICAL FORMULARY**



# The Chemical Formulary

*A Collection of Valuable, Timely, Practical  
Commercial Formulae and Recipes for  
Making Thousands of Products in  
Many Fields of Industry*

VOLUME IV

*Editor-in-Chief*

H. BENNETT



1939

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# **The Chemical Formulary, Volume IV**

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## P R E F A C E

Chemistry as taught in our schools and colleges is confined to synthesis, analysis and engineering—and properly so. It is part of the proper foundation for the education of the chemist.

Many a chemist on entering an industry soon finds that the bulk of the products manufactured by his concern are not synthetic or definite chemical compounds but are mixtures, blends or highly complex compounds of which he knows little or nothing. The literature in this field, if any, may be meagre, scattered or antiquated.

Even chemists, with years of experience in one or more industries, spend considerable time and effort in acquainting themselves on entering a new field. Consulting chemists, similarly, have problems brought to them from industries foreign to them. A definite need has existed for an up-to-date compilation of formulae for chemical compounding and treatment. Since the fields to be covered are many and varied, an editorial board was formed, composed of chemists and engineers in many industries.

Many publications, laboratories, manufacturing companies and individuals have been drawn upon to obtain the latest and best information. It is felt that the formulae given in this volume will save chemists and allied workers much time and effort.

Manufacturers and sellers of chemicals will find in these formulae new uses for their products. Non-chemical executives, professional men and others, who may be interested, will gain from this volume a "speaking acquaintance" with products which they may be using, trying, or with which they are in contact.

It often happens that two individuals using the same ingredients in the same formula get different results. This may be the result of slight deviations or unfamiliarity with the intricacies of a new technique. Accordingly, repeated experiments may be necessary to get the best results. Although many of the formulae given are being used commercially many have been taken from patent specifications and the literature. Since these sources are often subject to various errors and omissions, due regard must be given to this factor. Wherever possible it is advisable to consult with other chemists or technical workers regarding commercial production. This will save time and money and avoid "headaches."

It is seldom that any formula will give exactly the results which one requires. Formulae are useful as starting points from which to work out one's own ideas. Formulae very often give us ideas which may help us in our specific problems. In a compilation of this kind errors of omission, commission and printing may occur. We shall be glad to receive any constructive criticism in this, our first attempt.

To the layman, it is suggested that he arrange for the services of a chemist or technical worker familiar with the specific field in which he is interested. Although this involves an expense it will insure quicker and better formulation without wastage of time and materials.

H. BENNETT



## P R E F A C E   T O   V O L U M E   I V

New and additional formulae gathered during the past year have accumulated in quantity sufficient to produce an additional volume to volumes I, II & III of the Chemical Formulary. With some trepidation as to the acceptability of a fourth volume at this time, inquiry was made of educators, chemists, engineers, and manufacturers as to whether or not it should be published at this time. The unanimous answer was that information of this nature should be disseminated at the earliest moment to do the greatest amount of good.

It is gratifying to note that schools and colleges are making greater use of the Chemical Formulary as an auxiliary in stimulating practical interest in chemistry. With it students are making adhesives, insecticides, polishes, cosmetics, etc. It makes chemistry interesting, practical and useful to them.

Because some purchasers of this book are beginners in the art of chemical compounding, it was suggested that a simple introductory chapter be included. This has been done, giving complete directions for making simple preparations in everyday use.

It is a sincere pleasure to acknowledge the valuable assistance of the members of the board of editors and others who have given of their time and knowledge in contributing the special formulae which have made this volume possible.

H. BENNETT



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## A B B R E V I A T I O N S

amp.	ampere
amp./dm <sup>2</sup>	ampères per square decimeter
amp./sq. ft.	ampères per square foot
anhydr.	anhydrous
avoir.	avoirdupois
Bé.	Baumé
b.p.	boiling point
C.	Centigrade
°C.	Degrees Centigrade
cc.	cubic centimeter
c.d.	current density
cm.	centimeter
cm <sup>3</sup>	cubic centimeter
conc.	concentrated
c.p.	chemically pure
cu. ft.	cubic foot
cu. in.	cubic inch
cwt.	hundred weight
d.	density
dil.	dilute
dm.	decimeter
dm <sup>2</sup>	square decimeter
dr.	dram
E.	Engler
F.	Fahrenheit
°F.	Degrees Fahrenheit
f.f.c.	free from chlorine
f.f.p.a.	free from prussic acid
f. dr.	fluid dram
f. oz.	fluid ounce
f.p.	freezing point
ft.	foot
ft. <sup>2</sup>	square foot
g.	gram
gal.	gallon
gr.	grain
hl.	Hectoliter
hr.	hour
in.	inch
kg.	kilogram
l.	liter
lb.	pound
liq.	liquid
m.	meter
min.	minim, minute
ml.	milliliter—cubic centimeter
mm.	millimeter
m.p.	melting point
N.	Normal
N.F.	National Formulary
oz.	ounce
pH	Hydrogen-Ion Concentration

## ABBREVIATIONS

---

p.p.m.	parts per million
pt.	pint
pwt.	pennyweight
q.s.	a quantity sufficient to make
qt.	quart
r.p.m.	revolutions per minute
S.A.E.	Society of Automotive Engineers
sec.	second
sp.	spirits
sp. gr.	specific gravity
sq. dm.	square decimeter
tech.	technical
tinc.	tincture
tr.	tincture
Tw.	Twaddell
U.S.P.	United States Pharmacopeia
v.	volt
visc.	viscosity
vol.	volume
wt.	weight

## INTRODUCTION

At the suggestion of a number of teachers of chemistry and home economics the following introductory matter has been included.

The contents of this section are written in a simple way so that anyone, regardless of technical education or experience, can start making simple products without any complicated or expensive machinery. For commercial productions, however, suitable equipment is necessary.

Chemical specialties en masse are composed of pigments, gums, resins, solvents, oils, greases, fats, waxes, emulsifying agents, water, chemicals of great diversity, dyestuffs, and perfumes. To compound certain of these with some of the others requires certain definite and well-studied procedure, any departure from which will inevitably result in failure. The successful steps are given with the formulas. Follow them explicitly. If the directions require that A should be added to B, carry this out literally, and not in reverse fashion. In making an emulsion, the job is often quite as tricky as the making of mayonnaise. In making mayonnaise, you add the oil to the egg, *slowly*, with constant and even and regular stirring. If you do it correctly, you get mayonnaise. If you depart from any of these details: if you add the egg to the oil, or pour the oil in too quickly, or fail to stir regularly, the result is a complete disappointment. The same disappointment might be expected if the prescribed procedure of any other formula is violated.

The next point in importance is the scrupulous use of the proper ingredients. Substitutions are sure to result in inferior quality, if not in complete failure. Use what the formula calls for. If a cheaper product is desired, do not obtain it by substituting a cheaper material for the one prescribed: resort to a different formula. Not infrequently a formula will call for some ingredient which is difficult to obtain: in such cases, either reject the formula or substitute a similar material only after preliminary experiment demonstrates its usability. There is a limit to which this rule may reasonably be extended. In some instances the substitution of an equivalent ingredient may legitimately be made. For example: when the formula calls for *white wax* (beeswax), yellow wax can be used, if the color of the finished product is a matter of secondary importance. Yellow beeswax can often replace white beeswax, making due allowance for color: but paraffin will *not* replace beeswax, even though its light color recommends it above yellow beeswax.

And this leads to the third point: the use of good quality ingredients, and ingredients of the correct quality. Ordinary lanolin is not the same thing as *anhydrous* lanolin: the replacement of one for the other, weight for weight, will give discouragingly different results. Use exactly what the formula calls for: if you are unacquainted with the material and a doubt arises as to just what is meant, discard the formula and use one that you understand. Buy your materials from reliable sources. Many ingredients are obtainable in a number of different grades: if the formula does not designate the grade, it is understood that the best grade is to be used. Remember that a formula and the directions can tell you only a part of the story. Some skill is often required to attain success. Practice with a small batch in such cases until you are sure of your technique. Many instances can be cited. If the formula calls for steeping quince seed for 30 minutes in cold water, your duplication of this procedure may produce a mucilage of too thin a consistency. The originator of the formula may have used a fresher grade of seed, or his conception of what "cold" water means may be different from yours. You should have a feeling for the right degree of mucilaginousness, and if steeping the seed for 30 minutes fails to produce it, steep them longer until you get the right kind of mucilage. If you do not know what the right kind is, you will have to experiment until you find out. Hence the recommendation to make small experimental batches until successful results are arrived at. Another case is the use of dyestuffs for coloring lotions, and the like. Dyes vary in strength: they are all very powerful in tinting value: it is not always easy to state in quantitative terms how much to use. You must establish the quantity by carefully adding minute quantities until you have the desired tint.

## INTRODUCTION

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Gum tragacanth is one of those products which can give much trouble. It varies widely in solubility and bodying power; the quantity prescribed in the formula may be entirely unsuitable for your grade of tragacanth. Hence a correction is necessary, which can only be made after experiments to determine *how much* to correct.

In short, if you are completely inexperienced, you can profit greatly by gaining some experience through recourse to experiment. Such products as mouth washes, hair tonics, astringent lotions, need little or no experience, because they are as a rule merely mixtures of simple liquid and solid ingredients, the latter dissolving without difficulty and the whole being a clear solution that is ready for use when mixed. On the other hand, face creams, tooth pastes, lubricating greases, wax polishes, etc., which require relatively elaborate procedure and which depend for their usability on a definite final viscosity, must be made with the exercise of some skill, and not infrequently some experience.

### *Figuring*

Some prefer proportions expressed by weight, volume or in terms of percentages. In different industries and foreign countries various systems of weights and measures are used. For this reason no one set of units could be satisfactory for everyone. Thus divers formulae appear with different units in accordance with their sources of origin. In some cases, parts instead of percentages or weight or volume is designated. The following examples illustrate typical units:

#### Ink for Marking Glass

Glycerin	40	Ammonium Sulphate	10
Barium Sulphate	15	Oxalic Acid	8
Ammonium Bifluoride	15	Water	12

Here no units are mentioned. When such is the case it is standard practice to use parts by weight, using the same system throughout. Thus here we may use ounces or grams as desired. But if ounces are used for one item then ounces must be the unit for all the other items in the particular formula.

#### Flexible Glue

Glue, Powdered	30.9 %	Glycerin	5.15%
Sorbitol (85%)	15.45%	Water	48.5 %

Where no units of weight or volume but percentages are given then forget the percentages and use the same instructions as given under Example No. 1.

#### Example No. 3

#### Antiseptic Ointment

Petrolatum	16 parts	Benzoic Acid	1 part
Coconut Oil	12 parts	Chlorthymol	1 part
Salicylic Acid	1 part		

The same instructions as given under Example No. 1 apply to Example No. 3.

It is not wise in many cases to make up too large a quantity of material until one has first made a number of small batches to first master the necessary technique and also to see whether it is suitable for the particular outlet for which it is intended. Since, in many cases, a formula may be given in proportions as made up on a commercial factory scale, it is advisable to reduce the proportions accordingly. Thus, taking the following formula:

#### Example No. 4

#### Neutral Cleansing Cream

Mineral Oil	80 lb.	Water	90 lb.
Spermaceti	30 lb.	Glycerin	10 lb.
Glyceryl Monostearate	24 lb.	Perfume	to suit

Here, instead of pounds, grams may be used. Thus this formula would then read:

Mineral Oil	80 g.	Water	90 g.
Spermaceti	30 g.	Glycerin	10 g.
Glyceryl Monostearate	24 g.	Perfume	to suit

Reduction in bulk may also be obtained by taking the same fractional part or portion of each ingredient in a formula. Thus in the following formula:

**Example No. 5**

Vinegar Face Lotion			
Acetic Acid (80%)	20	Alcohol	440
Glycerin	20	Water	500
Perfume	20		

We can divide each amount by ten and the finished bulk is only 1/10th of the original formula. Thus it becomes:

Acetic Acid (80%)	2	Alcohol	44
Glycerin	2	Water	50
Perfume	2		

*Apparatus*

For most preparations pots, pans, china and glassware, such as is used in every household, will be satisfactory. For making fine mixtures and emulsions a "malted-milk" mixer or egg-beater is necessary. For weighing, a small, low priced scale should be purchased from a laboratory supply house. For measuring of fluids, glass graduates or measuring glasses may be purchased from your local druggist. Where a thermometer is necessary a chemical thermometer should be obtained from a druggist or chemical supply house.

*Methods*

To better understand the products which you intend making, it is advisable that you read the complete section covering such products. Very often an important idea is thus gotten. You may learn different methods that may be used and also avoid errors which many beginners are prone to make.

*Containers for Compounding*

Where discoloration or contamination is to be avoided (as in light colored, or food and drug products) it is best to use enamelled or earthenware vessels. Aluminum, as well, is highly desirable in such cases but it should not be used with alkalies as the latter dissolve and corrode this metal.

*Heating*

To avoid overheating, it is advisable to use a double boiler when temperatures below 212° F. (temperature of boiling water) will suffice. If a double boiler is not at hand, any pot may be filled with water and the vessel containing the ingredients to be heated is placed therein. The pot may then be heated by any flame without fear of overheating. The water in the pot, however, should be replenished from time to time as necessary—it must not be allowed to "go dry." To get uniform higher temperatures, oil, grease or wax is used in the outer container in place of water. Here of course care must be taken to stop heating when thick fumes are given off as these are inflammable. When higher uniform temperatures are necessary, molten lead may be used as a heating medium. Of course, where materials melt uniformly and stirring is possible, direct heating over an open flame is possible.

Where instructions indicate working at a certain temperature, it is important that the proper temperature be attained—not by guesswork, but by the use of a thermometer. Deviations from indicated temperatures will usually result in spoiled preparations.

*Temperature Measurements*

In Great Britain and the United States, the Fahrenheit scale of temperature measurement is used. The temperature of boiling water is 212° Fahrenheit (212° F.); the temperature of melting ice is 32° Fahrenheit (32° F.).

In scientific work and in most foreign countries the Centigrade scale is used. On this scale of temperature measurement, the temperature of boiling water is 100 degrees Centigrade (100° C.) and the temperature of melting ice is 0 degrees Centigrade (0° C.).

The temperature of liquids is measured by a glass thermometer. The latter is inserted as deeply as possible in the liquid and is moved about until the temperature remains steady. It takes a little time for the glass of the thermometer to come to

## INTRODUCTION

the temperatures of the liquid. The thermometer should not be placed against the bottom or side of the container, but near the center of the liquid in the vessel. Since the glass of the bulb of the thermometer is very thin, it can be broken easily by striking it against any hard surface. A cold thermometer should be warmed gradually (by holding over the surface of a hot liquid) before immersion. Similarly the hot thermometer when taken out should not be put into cold water suddenly. A sharp change in temperature will often crack the glass.

### *Mixing and Dissolving*

Ordinary solution (e.g. sugar in water) is hastened by stirring and warming. Where the ingredients are not corrosive, a clean stick, bone or composition fork or spoon is used as a mixing device. These may also be used for mixing thick creams or pastes. In cases where most efficient stirring is necessary (as in making mayonnaise, milky polishes, etc.) an egg beater or a malted-milk mixer is necessary.

### *Filtering and Clarification*

When dirt or undissolved particles are present in a liquid, they are removed by settling or filtering. In the former the solution is allowed to stand and if the particles are heavier than the liquid they will gradually sink to the bottom. The upper liquid may be poured or siphoned off carefully and in some cases is then of sufficient clarity to be used. If, however, the particles do not settle out then they must be filtered off. If the particles are coarse they may be filtered or strained through muslin or other cloth. If they are very small particles then filter paper is used. Filter papers may be obtained in various degrees of fineness. Coarse filter paper filters rapidly but will not, of course, take out extremely fine particles. For the latter, it is necessary to use a very fine grade of filter paper. In extreme cases even this paper may not be fine enough. Here it will be necessary to add to the liquid 1-3% of infusorial earth or magnesium carbonate. The latter clog up the pores of the filter paper and thus reduce their size and hold back undissolved material of extreme fineness. In all such filtering, it is advisable to take the first portions of the filtered liquid and pour them through the filter again as they may develop cloudiness in standing.

### *Decolorizing*

The most commonly used decolorizer is decolorizing carbon. The latter is added to the liquid to the extent of 1-5% and heated with stirring for  $\frac{1}{2}$  hour to as high a temperature as is feasible. It is then allowed to stand for a while and filtered. In some cases bleaching must be resorted to. Examples of this are given in this book.

### *Pulverizing and Grinding*

Large masses or lumps are first broken up by wrapping in a clean cloth and placing between two boards and pounding with a hammer. The smaller pieces are then pounded again to reduce their size. Finer grinding is done in a mortar with a pestle.

### *Spoilage and Loss*

All containers should be closed when not in use to prevent evaporation or contamination by dust; also because, in some cases, air affects the material adversely. Many materials attack or corrode the metal containers in which they are received. This is particularly true of liquids. The latter, therefore, should be transferred to glass bottles which should be as full as possible. Corks should be covered with aluminum foil (or dipped in melted paraffin wax when alkalies are present).

Materials such as glue, gums, olive oil or other vegetable or animal products may ferment or become rancid. This produces discoloration or unpleasant odors. To avoid this, suitable antiseptics or preservatives must be used. Too great stress cannot be placed on cleanliness. All containers must be cleaned thoroughly before use to avoid various complications.

### *Weighing and Measuring*

Since, in most cases, small quantities are to be weighed, it is necessary to get a light scale. Heavy scales should not be used for weighing small amounts as they are not accurate for this type of weighing.

For measuring volume (liquids) measuring glasses or cylinders (graduates) should be used. Since this glassware cracks when heated or cooled suddenly it should not be subjected to sudden changes of temperature.

*Caution*

Some chemicals are corrosive and poisonous. In many cases they are labeled as such. As a precautionary measure, it is advised not to smell bottles directly, but only to sniff a few inches from the cork or stopper. Always work in a well ventilated room when handling poisonous or unknown chemicals. If anything is spilled, it should be wiped off and washed away at once.

*Where to Buy Chemicals and Apparatus*

Many chemicals and most glassware can be purchased from your druggist. Notices of suppliers of all products will be found at the end of this book.

*ADVICE*

This book is the result of co-operation of many chemists and engineers who have given freely of their time and knowledge. It is their business to act as consultants and, for a fee, to give advice on technical matters. As publishers, we do not maintain a laboratory or consulting service to compete with them.

Please, therefore, do not ask us for advice or opinions, but confer with a chemist in your vicinity.

*Extra Reading*

Keep up with new developments of new materials and methods by reading technical magazines. Many technical publications are listed under references in the back section of this book.

*Calculating Costs*

Purchases of raw materials, in small quantities, are naturally higher in price than when bought in large quantities. Commercial prices, as given in the trade papers and catalogs of manufacturers, are for quantities such as barrels, drums or sacks. For example, a pound of epsom salts, bought at retail, may cost 10 or 15 cents. In barrel lots its price today is about 2 to 3 cents per pound.

Typical Costing Calculation		
Formula for Beer- or Milk Pipe Cleaner		
Soda Ash	25 lb. @ .02½ per lb. =	\$0.63
Sodium Perborate	75 lb. @ .16 per lb. =	12.00
Total 100 lb.		Total \$12.63

If 100 lb. cost \$12.63, 1 lb. will cost \$12.63 divided by 100 or about \$0.126 per lb. for raw materials, assuming no loss.

Always weigh the amount of finished product and use *this* weight in calculating costs. Most compounding results in some loss of material because of spillage, sticking to apparatus, evaporation, etc. Costs of making experimental lots are always high and should not be used for figuring costs. To meet competition, it is necessary to buy in larger units and costs should be based on the latter.

*Elementary Preparations*

The recipes that follow have been gotten up in a very simple way. Only one of each type is given so as to avoid confusion. These have been selected because of their importance and because they can be made readily.

The succeeding chapters go into greater detail and give many different types and modifications of these and other recipes for home and commercial use.

*Cleansing Creams*

Cleansing creams as the name implies serve as skin cleaners. Their basic ingredients are oils and waxes which are rubbed into the skin. When wiped off they carry off dirt and dead skin. The liquefying type of cleansing cream con-

tains no water and melts or liquefies when rubbed on the skin. To suit different climates and likes and dislikes harder or softer products can be made.

*Cleansing Cream (Liquefying)*

Liquid Petrolatum (White Mineral Oil)	5½ oz.
Paraffin Wax	2½ oz.
Petrolatum (Vaseline)	2 oz.
Melt together with stirring in an aluminum or enamelled dish and allow to cool. Then stir in a dash of perfume oil. Allow to stand until a haziness appears and then pour into jars, which should be allowed to stand <i>undisturbed</i> over night.	

*Cold Creams*

The most important facial cream is cold cream. This type of cream consists

of a mineral oil and wax which are emulsified in water with a little borax or glycostein. The function of a cold cream is to furnish a greasy film which takes up dirt and waste tissue which are removed when the skin is wiped thoroughly. Many modifications of this basic cream are encountered in stores. They vary in color, odor, and in claims but, essentially, they are no more useful than this simple cream. The latest type of cold cream is the non-greasy cold cream which is of particular interest because it is non-alkaline and therefore non-irritating to sensitive skins.

**Cold Cream**

Liquid Petrolatum (White Mineral Oil)	52 g.
White Beeswax	14 g.

Heat the above in an aluminum or enamelled double boiler (the water in the outer pot should be brought to a boil). In a separate aluminum or enamelled pot dissolve

Borax	1 g.
Water	33 c.c.

and bring this to a boil. Add this in a thin stream, to the melted wax, while stirring vigorously in one direction only, to the melted wax mixture. Use a fork for stirring. When the mixture turns to a smooth thin cream, immerse the bottom of the thermometer in it from time to time, stirring continuously. When the temperature drops to 140° F. add ½ c.c. of perfume oil and continue stirring until the temperature drops to 120° F. At this point pour into jars where the cream will "set" after a while. If a harder cream is desired, reduce the amount of liquid petrolatum. If a softer cream is wanted increase it.

**Cold Cream (Non-Greasy)**

White Paraffin Wax	1 ¼ oz.
Petrolatum (Vaseline)	1 ½ oz.
Glycostein or Glyceryl	
Monostearate	2 ¼ oz.
Liquid Petrolatum (White Mineral Oil)	3 oz.

Heat the above in an aluminum or enamelled double boiler (the water in the outer pot should be boiling). Stir until clear. To this slowly add, while stirring vigorously with a fork,

Water (boiling)	10 oz.
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Continue stirring until smooth and then add with stirring, a little perfume oil. Pour into jars at 110–130° F. and cover the jars as soon as possible.

***Vanishing Creams***

Vanishing creams are non-greasy creams, soapy in nature. Some are white and others have a very beautiful pearly appearance. This type of cream depends on its soapiness for its cleansing character and is useful as a powder base.

**Vanishing Cream**

Stearic Acid 18 oz.  
Melt the above in an aluminum or enamelled double boiler (the water in the outer pot must be boiling). To the above add, in a thin stream, while stirring vigorously with a fork, the following boiling solution made in an aluminum or enamelled pot:

Potassium Carbonate	1/4 oz.
Glycerin	6 ½ oz.
Water	5 lb.

Continue stirring until the temperature falls to 135° F., then stir in a little perfume oil and stir from time to time until cold. Allow to stand over night and stir again the next day. Pack into jars which should be closed tightly.

***Hand Lotions***

Hand lotions are usually clear or milky liquids or salves which are useful in protecting the skin from roughness and redness because of exposure to cold, hot water, soap and other materials. "Chapped" hands are a common occurrence. The use of a good hand lotion keeps the skin smooth, soft, and in a normally healthy condition. The lotion is best applied at night, rather freely, and cotton gloves may be worn to prevent soiling. During the day it should be put on sparingly and the excess wiped off.

**Hand Lotion (Salve)**

Boric Acid	1 oz.
Glycerin	6 oz.

Warm the above in an aluminum or enamelled dish and stir with a clean wooden stick until dissolved (clear). Then allow to cool and work into the following mixture with a potato masher, or rounded stick, adding only a little of the above liquid at a time to the mixture below and not adding a further portion until it is fully absorbed.

Lanolin	6 oz.
Petrolatum or "Vaseline"	8 oz.

If it is desired to impart a pleasant odor to this lotion a little perfume may be added and worked in.

**Hand Lotion (Milky Liquid)**

Lanolin	1/4 teaspoonful
Glycostein or Glyceryl	

Monostearate	1 oz.
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Tincture of Benzoin 2 oz.  
Witch Hazel 25 oz.  
Melt the first two items together in an aluminum or enamelled double boiler. If no double boiler is at hand improvise one by standing the dish in a small pot containing boiling water. When the mixture becomes clear remove from the double boiler and add slowly, while stirring vigorously with a fork or stick, the tincture of benzoin and then the witch hazel. Continue stirring until cool and then put into one or two large bottles and shake vigorously. The finished lotion is a beautiful milky liquid comparable to the best hand lotions on the market sold at high prices.

*Brushless Shaving Creams*

Brushless or latherless shaving creams are soapy in nature and do not require lathering or water. The formula given below is of the latest type being free from alkali and non-irritating. It should be borne in mind, however, that certain beards are not softened by this type of cream and require the old-fashioned lathering shaving cream.

*Brushless Shaving Cream*

White Mineral Oil	10 oz.
Glycostein or Glyceryl	
Monostearate	10 oz.
Water	50 oz.

Heat the first two ingredients together in a pyrex or enamelled dish to 150° F. and into this run slowly, while stirring with a fork, the water which has been heated to boiling. Allow to cool to 105° F. and while stirring add a few drops of perfume oil. Continue stirring until cold.

*Mouth Washes*

Mouth washes and oral antiseptics are of practically negligible value. Many, however, insist on their use because of their refreshing taste and deodorizing value.

*Mouth Wash*

Benzoic Acid	5/8 oz.
Tincture of Rhatany	3 oz.
Alcohol	20 oz.
Peppermint Oil	1/8 oz.

Just shake together in a dry bottle until it is dissolved and it is ready. A teaspoonful is used to a small wine-glassful of water.

*Tooth Powders*

Tooth powders depend for their cleansing action on soap and mild abrasives such as precipitated chalk and magnesium carbonate. The antiseptic present is

practically of no value. The flavoring ingredients mask the taste of the soap and give the user's mouth a pleasant after-taste.

*Tooth Powder*

Magnesium Carbonate	420 g.
Precipitated Chalk	565 g.
Sodium Perborate	55 g.
Sodium Bicarbonate	45 g.
Soap, Powdered White	50 g.
Sugar, Powdered	90 g.
Wintergreen Oil	8 cc.
Cinnamon Oil	2 cc.
Menthol	1 g.

Dissolve the last three ingredients together and then rub well into the sugar. Add the soap and perborate mixing in well. Add the chalk with good mixing and then the sodium bicarbonate and magnesium carbonate. Mix thoroughly and sift through a fine wire screen. Keep dry.

*Foot Powders*

Foot powders consist of a filler such as talc or starch with or without an antiseptic or deodorizer. In the following formula the perborates liberate oxygen when in contact with perspiration which tends to destroy unpleasant odors. The talc acts as a lubricant and prevents friction and chafing.

*Foot Powder*

Sodium Perborate	3 oz.
Zinc Peroxide	2 oz.
Talc	15 oz.

Shake together thoroughly in a dry container until uniformly mixed. This powder must be kept dry or it will spoil.

*Liniments*

Liniments usually consist of an oil and an irritant such as methyl salicylate or turpentine. The oil acts as a solvent and tempering agent for the irritant. The irritant produces a rush of blood and warmth which is often slightly helpful.

*Liniment, Sore Muscle*

Olive Oil	6 fl. oz.
Methyl Salicylate	3 fl. oz.

Shake together and keep in a well stoppered bottle. Apply externally but do not apply to chafed or cut skin.

*Chest-Rubs*

In spite of the fact that chest-rubs are practically useless countless sufferers use them. Their action is similar to that of liniments and they differ only in that they are in the form of a salve.

## INTRODUCTION

**"Chest-Rub" Salve**

Yellow Petrolatum or	
Yellow Vaseline	1 lb.
Paraffin Wax	1 oz.
Eucalyptus Oil	2 fl. oz.
Menthol	½ oz.
Cassia Oil	¼ fl. oz.
Turpentine	½ fl. oz.

Melt the vaseline and paraffin wax together in a double boiler and then add the menthol. Remove from the heat, stir, and cool a little; then stir in the oils, turpentine, and acid. When it begins to thicken pour into tins and cover.

*Insect Repellants*

Preparations of this type may irritate sensitive skins. Moreover, they will not always work. Psychologically they often are helpful, even though they may not keep insects away, because they give one confidence of protection.

*Mosquito Repelling Oil*

Cedar Oil	2 fl. oz.
Citronella Oil	4 fl. oz.
Spirits of Camphor	8 fl. oz.

Just shake together in a dry bottle and it is ready for use. This preparation may be smeared on the skin as often as is necessary to repel mosquitoes and other insects.

*Fly Sprays*

Fly sprays usually consist of deodorized kerosene, perfuming material, and an active insecticide. In some cases they merely stun the flies who may later recover and begin buzzing again.

*Fly Spray*

Deodorized Kerosene	89 fl. oz.
Methyl Salicylate	1 fl. oz.
Pyrethrum Powder	10 oz.

Mix thoroughly by stirring from time to time; allow to stand covered over night and then filter through muslin.

*Caution:* This spray is inflammable and should not be used near open flames.

*Deodorant Spray*

(For public buildings, sick-rooms, laboratories, etc.)

Pine Needle Oil	2 oz.
Formaldehyde	2 oz.
* Acetone	6 oz.
* Isopropyl Alcohol	20 oz.

One ounce of the above is mixed with a pint of water for spraying.

\* Inflammable.

*Cresol Disinfectant*

* Caustic Soda	25½ g.
Water	140 cc.

\* Do not get this on skin as it is corrosive.

Dissolve the above in a pyrex or enamelled dish and warm it. To this add slowly the following warmed mixture:

* Cresylic Acid	500 cc.
Rosin	170 g.

Stir until dissolved and add water to make 1000 cc.

\*Do not get this on skin as it is corrosive.

*Ant Poison*

Sugar	1 lb.
Water	1 qt.
* Arsenate of Soda	125 g.
Boil and stir until uniform; strain through muslin; add a spoonful of honey.	

\* Poison.

*Bedbug Exterminator*

† Kerosene	90 fl. oz.
Clove Oil	5 fl. oz.
* Cresol	1 fl. oz.
Pine Oil	4 fl. oz.

Simply shake and bottle.

† Inflammable.

\* Corrosive to skin.

*Mothproofing Fluid (Non-Staining)*

Sodium Aluminum Silico-	
fluoride	½ oz.
Water	98 oz.
Glycerin	½ oz.
Sulfate (Wetting-Agent)	
Stir until dissolved.	

*Fly Paper*

Rosin	32 oz.
Rosin Oil	20 oz.
Castor Oil	8 oz.

Heat the above in an aluminum or enamelled pot on a gas stove with stirring until all the rosin has melted and dissolved. While hot pour on firm paper sheets of suitable size which has been brushed with soap water just before coating. Smooth out the coating with a long knife or piece of thin flat wood and allow to cool. If a heavier coating is desirable increase the amount of rosin used. Similarly a thinner coating is gotten by reducing the amount of rosin. The finished paper should be laid flat and not exposed to undue heat.

*Household Baking Powder*

Bicarbonate of Soda	28 oz.
Mono Calcium Phosphate	35 oz.
Corn Starch	27 oz.

Mix the above powders thoroughly in a dry can by shaking and rolling for a half hour. Pack into dry air-tight tins as moisture will cause lumping.

*Malted Milk Powder*

Malt Extract, Powdered	5 oz.
Skimmed Milk, Powdered	2 oz.
Sugar, Powdered	3 oz.

Mix thoroughly by shaking and rolling in a dry can. Pack in an air-tight container.	Lampblack 2 oz. Water 40 fl. oz. Stir this in a glass or porcelain dish until dissolved. Do not expose this to strong light or it will spoil. Finally pour into a brown glass bottle. In using these solutions wet the cloth with solution A and allow to dry. Then write on it with solution B using a quill pen.
Cocoa Malt Powder	Marking Crayon (Green)
Corn Sugar 55 oz. Malt, Powdered, Mild 19 oz. Skim Milk, Powdered 12½ oz. Cocoa 13 oz. Vanillin ½ oz. Salt, Powdered ¾ oz.	Ceresin 8 oz. Carnauba Wax 7 oz. Paraffin Wax 4 oz. Beeswax 1 oz. Talc 10 oz. Chrome Green 3 oz.
Mix thoroughly and then run through a fine wire sieve.	Melt the first four ingredients in any container and then add the last two slowly while stirring. Remove from the heat and continue stirring until thickening begins. Then pour into molds. If other color crayons are desired, other pigments may be used. For example for black, use carbon or bone-black; for blue, Prussian blue; for red, orange chrome yellow.
Sweet Cocoa Powder	Antique Coloring for Copper
Cocoa 17½ oz. Sugar, Powdered 32½ oz. Vanillin ¾ g.	Copper Nitrate 4 oz. Acetic Acid 1 oz. Water 2 oz.
Mix thoroughly and sift.	Dissolve by stirring together in a glass or porcelain vessel. Pack in glass bottles.
Pure Lemon Extract	To Use: Wet the copper to be colored and apply the above solution hot.
Lemon Oil U.S.P. 6½ fl. oz. Alcohol 12½ fl. oz. Shake together in a gallon jug till dissolved.	Blue-Black Finish on Steel
Artificial Vanilla Flavor	a. Place object in molten sodium nitrate (700-800° F.) for 2-3 minutes. Remove and allow to cool somewhat; wash in hot water; dry and oil with mineral or linseed oil.
Vanillin ¾ oz. Coumarin ¼ oz. Alcohol 2 pt.	b. Place in following solution for 15 minutes:
Stir the above in a glass or china pitcher until dissolved. Then stir in the following solution which has been made by stirring in another pitcher.	Copper Sulphate ½ oz. Iron Chloride 1 lb. Hydrochloric Acid 4 oz. Nitric Acid ½ oz. Water 1 gal.
Sugar 12 oz. Water 5¼ pt. Glycerin 1 pt.	Then allow to dry for several hours; place in above solution again for 15 min.; remove and dry for 10 hours. Place in boiling water for ½ hour; dry and scratch brush very lightly. Oil with mineral or linseed oil and wipe dry.
Color brown by adding sufficient "burnt" sugar coloring.	Rust Prevention Compound
Canary Bird Food	Lanolin 1 oz. * Naphtha 2 oz.
Yolk of Eggs, Dried and Chopped 2 oz.	Mix until dissolved.
Poppy Heads (Coarse Powder) 1 oz.	The metal to be protected is cleaned with a dry cloth and then coated with the above composition.
Cuttlefish Bone (Coarse Powder) 1 oz.	* Inflammable—keep away from flames.
Granulated Sugar 2 oz. Soda Crackers, Powdered 8 oz.	
Mix well together.	
Writing Ink (Blue-Black)	
Naphthol Blue Black 1 oz. Gum Arabic, Powdered ½ oz. Carbolic Acid ¼ oz. Water 1 gal.	
Stir together in a glass or enamelled vessel until dissolved.	
Laundry Marking Ink (Indelible)	
A. Soda Ash 1 oz. Gum Arabic, Powdered 1 oz. Water 10 fl. oz.	
Stir the above until dissolved.	
B. Silver Nitrate 4 oz. Gum Arabic, Powdered 4 oz.	

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